Improving Quality of Municipal Solid Waste Compost through Mg-Modified Zeolite

H. Taheri Soudejani, M. Heidarpour, M. Shayannejad, H. Kazemian, H. Shariatmadari, and M. Afyuni

ABSTRACT

Developing materials to improve the efficiency of fertilizers and increase their performance is one of the main areas of research in the agricultural science. To the best of our knowledge, this is the first time that Mg-modified zeolite is used for modification of composting process of urban organic solid waste. Natural zeolite was clinoptilolite with a particle size smaller than 250 μm modified with MgCl₂·6H₂O salt. Different ratios (i.e. 5, 10, and 15%, on a weight basis) of natural and/or Mg-modified zeolite were used to modify the Municipal Solid Waste (MSW) compost. Experimental results revealed that the pH of the final compost product with the Mg-modified zeolite was effectively buffered at 7.7. Furthermore, in comparison to the control treatment (without zeolite), the compost with natural and Mg-modified zeolite showed 32.13 and 41.25% reduction in electrical conductivity, respectively. Ammonium retention of the compost containing 15% natural and Mg-modified zeolite was increased by 64.51 and 110.10%, respectively. Metals mobility test of the matured composts also revealed that adding zeolite, especially the Mg-modified zeolite, could decrease the mobility of Cu, Pb, Zn, Fe, Ni, and Mn metals. The most decrease in the mobility of metals was observed in Cu, which was about 44.16 and 47.82%, in the compost with the natural and Mg-modified zeolite, respectively.

Keywords: Ammonium retention, Composting, Conductivity reduction, instead of Composting, Metals mobility.

INTRODUCTION

Composting is a clean and sustainable method to manage and recycle the Municipal Solid Waste (MSW) (Diaz et al., 2011; Lim et al., 2016). Composting increases the capacity of waste disposal on environment, and helps its use in the agriculture land, thereby decreasing the usage of chemical fertilizers (Zhang et al., 2006; Toscano et al., 2013; Lim et al., 2016). It seems that organic farming is growing because of the positive effects of the compost on physical, biological, and chemical properties of soil (Maftoun and Moshiri, 2008; Watteau and Villemin, 2011). Large amounts of the MSW compost are needed to provide enough Nitrogen (N) and other nutrients for plants. However, salts and heavy metals contained in the MSW composts can be harmful to plants, causing several unfavorable effects in the soil (Hargreaves et al., 2008; Gautam et al., 2010).

Nowadays, agricultural science professionals are working to develop eco-
friendly materials and methods to improve the performance and efficiency of fertilizers (Behin and Sadeghi, 2016; Gholamhoseini et al., 2013). Natural zeolites are being used due to their high adsorption and ion exchange ability, helping to improve the quality of fertilizers and decrease their environmental effects (Gholamhoseini, et al., 2013; Wang et al., 2014).

Zeolites are microporous crystalline aluminosilicate minerals that consist of three-dimensional frameworks of AlO₄ and SiO₄ tetrahedra (Costa et al., 2012). Voids and pores are the main characteristics of zeolite materials. The interconnected voids and pores are engaged by water molecules and cations. The internal surface area of these channels are informed to reach as much as several hundred square meters per gram of zeolite, causing zeolites as very effective ion exchangers (Ramesh and Reddy, 2011). Zeolites are commonly used as commercial adsorbents in a wide range of environmental processes (Zorpas and Inglezakis, 2009; Kazemian et al., 2012).

Many types of research have shown that the quality of MSW composting and sewage sludge can be improved by adding zeolite (Turan and Ergun, 2007; Stylianou et al., 2008). Zeolites have also been used to adsorb heavy metals, and ammonia (NH₃-N) from compost (Chiang et al., 2007; Turan and Ergun, 2007; Turan, 2008; Villasenor et al., 2011). Zorpas (2011) showed that 70% of Cd, 67% of Cr and Mn, 46% of Zn, 43% of Ni, 41% of Pb, 37% of Fe and 27% of Cu were taken up from clinoptilolite. Turan (2008) demonstrated that by enhancing the amount of natural zeolite from 5 to 10%, the salinity uptake efficiencies increased from 66.64 to 88.92%, respectively. The losses of NH₃-N must be controlled during composting to enhance the agronomic value of the compost and reduce atmospheric pollution (Villasenor et al., 2011). Turan and Ergun (2007) showed that the natural zeolite readily trapped excess N during the processes and reduced NH₃-N odors. The adsorbed ammonium (NH₃-N) by zeolite and its reuse in slow release fertilizers have been reported in different studies (Gholamhoseini et al., 2013; Wang et al., 2014; Chan et al., 2016). Using 21% zeolite during the composting of manure increased the NH₃-N from 0.08 to 0.21% (Gholamhoseini et al., 2013). During the composting process, 22 and 38% N was lost in zeolite applied row and row without zeolite, respectively (Joghan et al., 2010).

In many applications, different modification processes have been carried out to improve the sorption capacity of the natural zeolite. The conventional modification/regeneration methods often include treatments with acid and salts of alkaline metals (Watanabe et al., 2005; Huang et al., 2014). Furthermore, a Magnesium (Mg) modified zeolite triggers two reaction mechanisms simultaneously. These include adsorption and struvite crystallization. The Mg-modified zeolite can adsorb NH₃-N and release Mg²⁺, which is a source of struvite precipitant employed in the sorption of NH₃-N and phosphates. Adsorption of NH₃-N and phosphate from wastewater using the Mg-modified zeolite has shown higher efficiency, as compared to the natural zeolite (Huang et al., 2014). Chan et al. (2016) showed that adding Mg and P salts increased NH₃-N retention, pH, and salinity effectively. A compost modified with 10% zeolite showed the effective reduction of EC from 6.45 to 2.82 mS cm⁻¹.

Retention of NH₃-N in compost is important because of the nutritional value. To the best of our knowledge, the effect of Mg-modified zeolite on improving the compost properties is not studied thoroughly. Therefore, the main objective of this study was to investigate the effects of natural and Mg-modified zeolite on the quality of the MSW compost.

MATERIALS AND METHODS

Materials Preparation

This study was performed at the Isfahan Compost Factory, Isfahan, Iran (51°43'E,
This plant receives 1,200 tons d⁻¹ of solid waste gathered from the Isfahan City and its surrounding areas. The organic waste constitutes about 70% of the solid waste. To prepare the feed material for composting process, the following process was followed: First, a crusher hammer was used to reduce physical size of the solid waste. As the next step, non-compostable material was separated from MSW by a vibration process. Third, during the movement of waste from conveyor, ferrous metals was separated using a strong magnet. Then, during the course of moving solid waste in rotary hull, plastic impurities were separated. Finally, the organic waste was fed into the decomposition site. In this research, the raw material in the decomposition site was used as feed for composting process. Particle size of materials was prepared to be smaller than 60 mm. The organic wastes were mixed with the zeolites at different ratios and the composting process was performed. Natural and Mg-modified zeolite were used for the modification of the compost.

**Natural and Modified Zeolite**

Natural zeolite was clinoptilolite obtained from Semnan Province, Iran. Zeolite with a particle size smaller than 250 μm was chosen using ASTM standard sieves. The natural zeolite had the following chemical compositions (mass%): SiO₂ = 65.90, Al₂O₃ = 11.20, Na₂O = 2.10, K₂O = 2.31, CaO = 3.20, Fe₂O₃ = 1.25, MgO = 0.52, LOI = 11.89, and SiO₂/Al₂O₃ = 5.9 (Malekian et al., 2011).

Zeolite was modified by MgCl₂·6H₂O salt. Prior to the modification process, the natural zeolite was washed with distilled water to remove soluble impurities and dried at the ambient conditions for 36 hours. Subsequently, 15 kg dried zeolite was suspended in 150 L of a 2M (MgCl₂·6H₂O) solution decanted and stirred for five minutes every 30 min during 48 hours. Finally, the modified zeolite sample was dried at the ambient temperature. The X-Ray diffraction technique (XRD) was used to analyze modified zeolite. Some close resemblance was observed between the diffraction patterns of the natural and Mg-modified zeolite, suggesting that the structural integrity of zeolite remained intact after modification (Figure 1). The XRD pattern showed that the sample of natural and Mg-modified zeolite mainly consisted of clinoptilolite with an ideal formula of KNa₂Ca₂(Si₂₉Al₁₇)O₇₂·24H₂O and some quartz (Figure 1). The cation exchange capacity (CEC) of natural and Mg-modified zeolite were measured using ammonium acetate method (Page et al., 1982). Concentrations of some typical alkaline and alkaline earth cations of the natural and

![Figure 1. XRD patterns of the natural and modified zeolites.](image-url)
modified zeolite were extracted by an ammonium acetate solution (Huang et al., 2014) and measured by the atomic adsorption Spectrometry (AA 3030 Perkins Elmer), as reported in Table 1.

**Composting Setup and Operation**

Twenty-one lysimeters made of PVC with 40 cm inner diameter and 60 cm height were used. A randomized complete block design was employed with seven treatments and three replications. Different dosages of 5, 10, and 15% of the natural and Mg-modified zeolite were added to the MSW at the beginning of the composting process. The composts were mixed manually, once every other days, throughout the composting period. Treatments were labeled as follows: F1 presents the compost without zeolite. F2, F3 and F4 are the composts treated with 5, 10, and 15% natural zeolite, respectively; F5, F6 and F7 are the composts treated with 5, 10, and 15% Mg-modified zeolite, respectively.

The composting process was carried out for two months. The temperature inside the lysimeters was recorded every 4 days by a Testothermometer. Representative 150 g compost samples were taken every other two weeks to measure the moisture content, electrical conductivity, pH, total N, NH\textsubscript{4}–N and NO\textsubscript{3}–N. Samples were taken from three different parts of lysimeter and mixed together. Moisture content, pH and electrical conductivity were immediately measured in the wet samples. The moisture content was determined by drying at 105°C for 24 hours. During the composting period, the moisture of the composting lysimeters was measured every 2 days and kept at 40 to 60% by adding distilled water whenever needed. The pH and EC were measured by a digital pH-meter and a conductometer (AZ-8301 Cond), respectively, in a 1:10 water extract (Thompson et al., 2002). A 2-molar KCl solution was used to extract the N content of the samples. NH\textsubscript{4}+–N and NO\textsubscript{3}–N contents were determined using the steam distillation Kjeltc (Gerhardt-Germany) apparatus and spectrophotometer (Perkin Elmer 550S-Germany), respectively. The Total Nitrogen (TN) was also measured using Kjeldahl method (Vandenabeele et al., 1990). Carbon, germination index and mobility of the heavy metals were measured at the end of the period. The total C was measured by the Walkley–Black method after drying samples at 105°C (Nelson and Sommers, 1996). The metals mobility of the matured compost was measured by extracting with diethylene-Triamine-Pentaacetic Acid (DTPA) solution. The concentration of heavy metals was measured by the Atomic Adsorption Spectrometry (AA 3030 Perkins Elmer). The Germination Index (GI) was determined using the relative seed germination and the relative root elongation of garden cress. Seed germination and root length tests were done on water extracts by mechanically shaking the fresh samples at a solid to distilled water ratio of 1:10 (w/v dry weight basis) for one hour. Later, 0.5 mL of each extract was pipetted into a sterilized plastic petri dish lined with Whatman filter paper. Ten cress seeds were evenly distributed on the filter paper and incubated at 25°C in the dark for 48 hours. Treatments were assessed by measuring the length of the root radical of the composts.

<table>
<thead>
<tr>
<th>Cations</th>
<th>Natural zeolite (meq L\textsuperscript{–1})</th>
<th>Modified zeolite (meq L\textsuperscript{–1})</th>
</tr>
</thead>
<tbody>
<tr>
<td>Mg</td>
<td>0.7±0.05</td>
<td>11.7±0.72</td>
</tr>
<tr>
<td>Ca</td>
<td>2.9±0.12</td>
<td>1.8±0.05</td>
</tr>
<tr>
<td>Na</td>
<td>12.2±0.46</td>
<td>2.7±0.15</td>
</tr>
<tr>
<td>K</td>
<td>0.3±0.03</td>
<td>0.1±0.01</td>
</tr>
</tbody>
</table>

\( ^d \) Results are mean±SD of three determinations.
and counting the number of germinated seeds. The GI was determined by the following formula (Wong et al., 2001).

\[
\text{GI} = \frac{\text{Seed germination} \times \text{Root length of treatment}}{\text{Seed germination} \times \text{Root length of control}} \times 100
\]

Statistical Analyses

One-way Analysis Of Variance (ANOVA) was employed to test the significance at 1% level (LSD) by the SAS software (2002).

RESULTS AND DISCUSSION

Change in the pH Profile

Variation of pH during the composting process is illustrated in Figure 2. The pH value of MSW was 5.88 before composting. The pH value was increased, reaching a value close to 8 at the end of composting for all treatments, except F6 and F7 (Figure 2). Jumnoodoo and Mohee (2012) suggested that the optimal pH during the composting process was between 7.0 and 8.0, because pH in that range facilitated higher microbial activities. According to the analysis of variance, the type of zeolite (natural and modified) and percentage of zeolite in the applied compost were significant factors (P < 0.01) (Table 2). Increasing the zeolite dosage and using the modified zeolite helped the normalization of pH. The results also showed that the maximum and minimum pH of the matured compost was recorded for F1 (control) and F7 (15% Mg-modified zeolite) treatments, respectively (Figure 2). Zeolite could decrease the pH of the compost product because of retaining NH$_3$-N as a result of cation adsorption and ion exchange ability (Chiang et al., 2007; Stylianou et al., 2008; Turan, 2008). In this research, pH was reduced using Mg-modified zeolite, while Chan et al. (2016) did not observe any significant change in pH by using MgO, K$_2$HPO$_4$ salts and zeolite. This could be attributed to the fact that the buffer capacity of the phosphate salt is higher than that of chlorides (i.e. HCl is stronger than H$_3$PO$_4$). NH$_3$-N volatilization is dramatically increased once the pH is raised above 8.0, causing N losses and odor problems (Raphael and Velmourougane, 2011). The result showed that Mg-modified zeolite was more effective than natural forms for reduction of pH. It seems that due to higher absorption of NH$_4$-N in compost.

![Figure 2](image_url)

**Figure 2.** Changes in the pH of different treatments. Different letters in each composting time indicate significant differences between the values (P < 0.01) by LSD test. Results are mean ± SD (error bars) of three determinations. F symbols are defined under Table 3 and in the text.
Table 2. Mean comparison of the type and percent of zeolite at the end of the composting process.\(^a\)

<table>
<thead>
<tr>
<th>Factors</th>
<th>pH</th>
<th>EC (mS cm(^{-1}))</th>
<th>NH(_4)-N (mg kg(^{-1}))</th>
<th>NO(_3)-N (mg kg(^{-1}))</th>
<th>Total nitrogen (%)</th>
</tr>
</thead>
<tbody>
<tr>
<td>Type of zeolite</td>
<td></td>
<td></td>
<td></td>
<td></td>
<td></td>
</tr>
<tr>
<td>Natural</td>
<td>8.75(^a)</td>
<td>3.45(^a)</td>
<td>507.50(^b)</td>
<td>162.78(^a)</td>
<td>1.59(^a)</td>
</tr>
<tr>
<td>Modified</td>
<td>8.05(^b)</td>
<td>2.99(^b)</td>
<td>631.94(^a)</td>
<td>154.29(^a)</td>
<td>1.57(^a)</td>
</tr>
<tr>
<td>Percent of zeolite</td>
<td></td>
<td></td>
<td></td>
<td></td>
<td></td>
</tr>
<tr>
<td>5%</td>
<td>8.60(^a)</td>
<td>3.70(^a)</td>
<td>446.25(^b)</td>
<td>172.03(^a)</td>
<td>1.65(^a)</td>
</tr>
<tr>
<td>10%</td>
<td>8.33(^b)</td>
<td>3.15(^b)</td>
<td>586.25(^b)</td>
<td>156.92(^b)</td>
<td>1.58(^b)</td>
</tr>
<tr>
<td>15%</td>
<td>8.26(^b)</td>
<td>2.82(^c)</td>
<td>676.66(^c)</td>
<td>146.66(^b)</td>
<td>1.53(^b)</td>
</tr>
</tbody>
</table>

\(^a\) In each column, data with the same alphabet are not significantly different (P < 0.01).

containing Mg-modified zeolite (see section “Changes in NH\(_4\)-N”), less amount of NH\(_3\)-N was produced, so the pH was decreased more.

**Change in Electrical Conductivity (EC)**

As shown in Figure 3, the EC value was reduced with time in the composting process for all treatments. At the beginning of the composting, the EC value of MSW was 6.59 mS cm\(^{-1}\). During the initial phase of composting, EC was reduced. After this stage, there was a small change in EC. The acceptable level of EC for the mature compost is 4 mS cm\(^{-1}\) (Karak *et al.*, 2013). The final EC values in all composts, except the control treatment (5.1 mS cm\(^{-1}\)), were in an acceptable range for agricultural use (Figure 3). According to the experimental results, the type (natural and modified) and percentage of zeolite in the applied compost could have a significant effect on the EC of the compost (P < 0.01) (Table 2). As shown in Table 2, the average of EC values in the composts with natural and modified zeolite, at the end of the composting process, was 3.45 and 2.99 mS cm\(^{-1}\), respectively. The maximum and minimum EC of the matured compost was observed for F1 (i.e. control) and F7 (i.e. 15% Mg-modified) treatments, respectively (Figure 3). Increasing the zeolite dosage and using the modified zeolite significantly reduced the EC of the compost. Other studies have reported the same results for the natural zeolite (Chiang

![Figure 3. Changes of EC in different treatments during the composting process. Different letters in each composting time indicate significant differences between the values (P < 0.01) by LSD test. Results are mean± SD (error bars) of three determinations.](image)
et al., 2007; Turan and Ergun, 2007; Stylianou et al., 2008; Turan, 2008; Zhang and Sun, 2015; Chan et al., 2016). The natural zeolite can have negative charges balanced by freely moving cations with the positive charges. This provides an ideal trap for positive cations like potassium or NH$_4^-$, which are released when required by plants. Zeolites porous structure leads to a large specific surface helping to trap and exchange valuable nutrients. The natural zeolite adsorbs Na$^+$ and Cl$^-$ that enter the cavities, consequently, reducing the salinity of composts (Turan, 2008). Zeolite not only exchanges cations, but also adsorbs ions; thus, the EC of the compost is reduced. The decomposition process may either lead to organic forms that are neutral (not charged as ions), or gaseous products which escape the mixture (Stylianou et al., 2008). Slightly higher CEC of Mg-modified zeolite (141.3 meq 100 gr$^{-1}$) compared to natural zeolite (130.5 meq 100 gr$^{-1}$) can be attributed to the fact that modification process will facilitate access to some of zeolites cationic sites that are otherwise inaccessible, therefore, more cations will be exchanged.

### Changes in NH$_4^-$N

The results indicated that the amount and type of zeolite (i.e. natural and modified) had significant effects on retaining NH$_4^-$N in the compost (P< 0.01) (Table 2). Both natural and modified zeolites retained more NH$_4^-$N in comparison to F1 (control treatment), while the modified zeolite was more effective (Figure 4). Many researchers have reported that the retention of NH$_4^-$N and NH$_3^-$N by clinoptilolite could be due to ion exchange and adsorption processes, respectively (Villasenor et al., 2011; Gholamhoseini et al., 2013). Huang et al. (2014) used the Mg-modified zeolite for the retention of NH$_4^-$N in a wastewater treatment process. The higher retention involves two reaction mechanisms: ion exchange and struvite crystallization (Huang et al., 2014) (see equations 2 and 3). Struvite is a phosphate mineral with the formula: NH$_4$MgPO$_4$·6H$_2$O. The Mg-modified zeolite can adsorb NH$_4$-N and release Mg$^{2+}$ (Equation 2). The released Mg$^{2+}$ is a source of struvite formation to retain more NH$_4^-$N and phosphate (Equation 3).

\[\text{Zeolite} - \text{Mg}^{2+} + \text{NH}_4^+ \leftrightarrow \text{Mg}^{2+} + \text{Zeolite} - \text{NH}_4^+ \] (2)

\[\text{Mg}^{2+}(aq) + \text{NH}_4^+(aq) + \text{PO}_4^{3-}(aq) + 6\text{H}_2\text{O} \leftrightarrow \text{MgNH}_4\text{PO}_4 \cdot 6\text{H}_2\text{O} \downarrow \] (3)

Chan (2016) showed that Struvite formation, through the addition of MgO and

![Figure 4](https://example.com/figure4.png)

**Figure 4.** Changes of NH$_4^+$-N in different treatments during the composting process. Different letters in each composting time indicate significant differences between the values (P< 0.01) by LSD test. Results are mean±SD (error bars) of three determinations.

753
K₂HPO₄ during composting, significantly conserved NH₄-N. In this study, it was found that changes in the compost pH increased the NH₄-N retention in the case of Mg-modified zeolite as compared to the natural and control treatments. Previous studies showed the optimum pH of 7 (Maranon et al., 2006), 8 (Saltali et al., 2007), and 8.5 (Huang et al., 2014) for the adsorption of NH₄-N by zeolites. Huang et al. (2014) reported that at pH> 8.5 more NH₄-N was converted into NH₃-N forms, which could not be adsorbed by zeolite, resulting in the decline of the retention ratio of the TN. The results also revealed that the NH₄-N level was stabilized at the end of the composting process. The control treatment had a lower NH₄-N content than the other treatments (Figure 4), according to the procedure reported before (Villasenor et al., 2011).

Changes in NO₃-N

The amount of NO₃-N was increased during the composting process (Figure 5). Ko (2008) showed that NH₄-N concentrations decreased during the composting process, while NO₃-N concentrations increased sharply after 30 days. Our results indicated that the amount of zeolite had a significant effect on the amount of compost NO₃-N (P< 0.01). More zeolite dosage caused more reduction in NO₃-N production (Table 2). These results were consistent with the previous studies (Gholamhoseini et al., 2013). A possible explanation could be that zeolite prevented nitrification and conversion of NH₄-N to NO₃-N (Taheri-Sodejani et al., 2015). Minimum and maximum NO₃-N levels at the end of process were observed in F7 and the control, respectively (Figure 5).

Changes in Total Nitrogen (TN), Total Weight of Compost, and C/N Ratio

TN contents during the composting process are shown in Figure 6. The results showed that the most reduction in the TN content occurred in the control treatment. Increasing the zeolite dosage caused less reduction in the TN content of the final compost. It has been reported that NH₃-N content causes reduction in the N content and some undesirable odor that could be increased during the composting process due to the anaerobic conditions at the higher pH levels (Turan and Ergun, 2007; Wang et al., 2014). Turan and Ergun (2007), Turan (2008), used 5 and 10% of zeolite, achieving the TN losses reduction of 11 and 9%, respectively. The compost quality with respect to the agricultural use depends on its inorganic N content, which should not exceed 10% of the TN (Zorpas et al., 1999; Zorpas et al.,...
2003). In this research, the total inorganic N did not exceed the limit, showing that the zeolite-modified compost produced in this research had an appropriate quality for agricultural use.

The total weight of the produced compost and N in different treatments is summarized in Table 3. The differences between treatments were significant at the 1% level. The highest amounts of compost and N were obtained in F4 and F7 treatments. Increasing the zeolite dosage enhanced the amount of the compost and the total N retained in the compost (Table 4). Retention of N in the compost is important for plants because of its nutritional value.

The carbon (C) content in different treatments is shown in Table 3. The ideal mature compost should have a total C content of 8 to 35% (Turan and Ergun, 2007). Therefore, the total C contents of the matured composts can be in the acceptable range. Increasing the zeolite dosage in the composting process caused the reduction in the C content due to the bulking effect (Table 4). The C/N ratio representing the compost maturity was decreased in all zeolite treatments. The C/N range of 9 to 10 could be considered suitable in the mature compost (Bernai et al., 1998; Zorbas et al., 1999). With increasing the rate of zeolite, C/N got close to the acceptable range.

**Germination Index (GI)**

The GI for the zeolite treatments was higher, as compared to the control treatment. This index was 100 for the control treatment, while it was 104.32 and 104.94 for the compost with the natural and modified zeolites, respectively.

### Table 3. Properties of the compost product after 60 days of composting.

<table>
<thead>
<tr>
<th>Treatments</th>
<th>Carbon (C) (%)</th>
<th>C/N</th>
<th>Amount compost product (kg)</th>
<th>Total N retained (gr)</th>
<th>Germination Index</th>
</tr>
</thead>
<tbody>
<tr>
<td>F1</td>
<td>24.60</td>
<td>13.14</td>
<td>3.41</td>
<td>63.88</td>
<td>100.00</td>
</tr>
<tr>
<td>F2</td>
<td>21.55</td>
<td>12.74</td>
<td>4.20</td>
<td>71.16</td>
<td>105.55</td>
</tr>
<tr>
<td>F3</td>
<td>17.81</td>
<td>11.46</td>
<td>5.34</td>
<td>84.48</td>
<td>103.70</td>
</tr>
<tr>
<td>F4</td>
<td>16.13</td>
<td>10.60</td>
<td>6.32</td>
<td>96.23</td>
<td>103.70</td>
</tr>
<tr>
<td>F5</td>
<td>21.05</td>
<td>13.08</td>
<td>5.85</td>
<td>74.24</td>
<td>105.55</td>
</tr>
<tr>
<td>F6</td>
<td>18.10</td>
<td>11.46</td>
<td>5.97</td>
<td>94.31</td>
<td>103.70</td>
</tr>
<tr>
<td>F7</td>
<td>15.54</td>
<td>10.04</td>
<td>6.92</td>
<td>106.90</td>
<td>105.55</td>
</tr>
</tbody>
</table>

*In each column, data with the same alphabet are not significantly different at P< 0.01.

*F1: Presents the compost without zeolite, and F2, F3 and F4 are composts treated with 5, 10, and 15% natural zeolite, respectively; F5, F6 and F7 are composts treated with 5, 10, and 15% Mg-modified zeolite, respectively.*

![Figure 6](#) Changes of the total N in different treatments during the composting process. Different letters in each composting time indicate significant differences between the values (P< 0.01) by LSD test. Results are mean± SD (error bars) of three determinations.
Table 4. Results of the mean comparison test of the type and percent of the zeolite effect.

<table>
<thead>
<tr>
<th>Factors</th>
<th>The amount of compost produced (Kg)</th>
<th>Total N retained</th>
<th>Carbon (C) (%)</th>
<th>C/N</th>
</tr>
</thead>
<tbody>
<tr>
<td></td>
<td></td>
<td></td>
<td></td>
<td></td>
</tr>
<tr>
<td>Type of zeolite</td>
<td></td>
<td></td>
<td></td>
<td></td>
</tr>
<tr>
<td>Natural</td>
<td>5.29b</td>
<td>83.96b</td>
<td>18.5a</td>
<td>11.53a</td>
</tr>
<tr>
<td>Modified</td>
<td>5.92a</td>
<td>93.15b</td>
<td>18.2a</td>
<td>11.53a</td>
</tr>
<tr>
<td>Percent of zeolite</td>
<td></td>
<td></td>
<td></td>
<td></td>
</tr>
<tr>
<td>5%</td>
<td>4.53c</td>
<td>74.70c</td>
<td>21.30c</td>
<td>12.91c</td>
</tr>
<tr>
<td>10%</td>
<td>5.65b</td>
<td>89.39b</td>
<td>17.95b</td>
<td>11.36b</td>
</tr>
<tr>
<td>15%</td>
<td>6.62a</td>
<td>101.56a</td>
<td>18.84a</td>
<td>10.33c</td>
</tr>
</tbody>
</table>

<i>Table 4. Results of the mean comparison test of the type and percent of the zeolite effect.</i>

<i>a In each column, data with the same alphabet are not significantly different at P < 0.01.</i>
Figure 7. Concentration of the heavy metals in the compost at the end of the composting process (mg kg⁻¹). F symbols are defined under Table 3 and in the text. [DL: Detection Limit; DL(Pb)=0.2 ppm; DL(Ni)=0.1 ppm; DL(Zn)=0.1 ppm; DL(Cu)=0.2 ppm; DL(Mn)=0.2 ppm; DL(Fe)=0.15 ppm].

The most decrease in the mobility of metals was observed in Cu, which was 44.16 and 47.82%, in the compost with natural and modified zeolite, respectively. Considering the experimental data of this study, we suggest to use natural clinoptilolite zeolite and its Mg-modified form to improve composting of MSW. Zeolite-modified compost should be considered as a value-added soil amender with higher content of nutrient that is produced via an environmental friendly process. Further
ACKNOWLEDGEMENTS

This work was supported by Waste Management Organization in Isfahan, Iran. Staff of this organization are gratefully acknowledged.

REFERENCES

بهبود خصوصیات کمپوست زباله شهری با کاربرد زئولیت اصلاح شده با منیسیم

م. ظاهری سودجانی، م. حیدرپور، م. شایان نژاد، ح. کاظمیان، ح. شریعتمداری، و م. افیونی

چکیده

توسعه مواد جدید برای بهبود راندمان استفاده از کوپَست و آنفایش عمکرد آن ها یکی از اهداف اصلی محققان در بخش کشاورزی است. مطالعات انجام شده نشان می‌دهد که این تحقیق برای اولین بار اثر زئولیت اصلاح شده با کمپوست زباله شهری بررسی می‌کند. زئولیت طبیعی آماده از نوع کلیپتیولیت و با اندازه ذرات کمتر از 25 میکرون بود که با استفاده از نمک منیسیم کلیسکارید اصلاح شد. مقادیر مختلفی از زئولیت طبیعی و اصلاح شده (5، 10 و 15 درصد وزنی) به منظور اصلاح کمپوست مورد استفاده قرار گرفت. نتایج مطالعه نشان داد که اکنون کمپوست با کاربرد کمپوست حاصل از زئولیت اصلاح شده pH استفاده قرار گرفته. نتایج مطالعه نشان داد که اکنون کمپوست حاصل از زئولیت اصلاح شده pH استفاده قرار گرفته. نتایج مطالعه نشان داد که اکنون کمپوست حاصل از زئولیت اصلاح شده pH استفاده قرار گرفته. نتایج مطالعه نشان داد که اکنون کمپوست حاصل از زئولیت اصلاح شده pH استفاده قرار گرفته.