

## Chemical Immobilization of Lead (Pb) in Long Term Sewage Irrigated Soil

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### ABSTRACT

Sewage effluent is the carrier of huge amounts of N (0.09 to 0.16%), P (0.023 to 0.029 %), and K (0.11 to 0.18%) and as well as it is rich in organic carbon (3.27 %). Transfer of the toxic trace metal (Pb) from the contaminated site can be seized out from the food chain through application of  $\text{FeSO}_4 \cdot 7\text{H}_2\text{O}$  and  $\text{CaSO}_4$  in a larger area. The efficacy of the toxic metal immobilization through an application of  $\text{FeSO}_4 \cdot 7\text{H}_2\text{O}$  and  $\text{CaSO}_4$  was studied in the laboratory and tested in the field. In batch experiments, application of  $\text{FeSO}_4$  at the rate of 5% (on soil weight basis) indicated a greater decrease in soil pH (from 7.54 to 5.55) in half an hour of shaking period as a result of which water soluble and exchangeable Pb concentrations decreased by 100%. Application of  $\text{CaSO}_4$  also decreases soil pH, as well as Pb concentration in soil solution. Pot culture experiments with two flower and leafy vegetable plants also show decrease in bioavailability of Pb in  $\text{FeSO}_4$  treated soils.

**Keywords:** Chemical amendments, Long term irrigation, Pb accumulations, Sewage.

### INTRODUCTION

The use of urban wastewater in agriculture is a common practice for diverse reasons, not least of which being water scarcity, fertilizers' costs, and many a times a lack of an alternative source of irrigation water (Raschid *et al.*, 2006). Sewage, often untreated, used to irrigate around 10% of the world's crops, especially in urban areas, reveals the first global survey of the hidden practice of wastewater irrigation (Scott *et al.*, 2004). More than 20 million ha of land are irrigated with urban wastewater. Wastewater has an important impact on agricultural productivity and livelihoods. Strauss and Blumenthal (1990) estimated that 73,000 ha of agricultural land were

irrigated with wastewater in India. A consideration of the potential availability of sewage revealed that 2,600 million  $\text{m}^3$  of sewage is produced per annum in India (Arti Bhatia *et al.*, 2001). Most irrigation through wastewater in India occurs along the river banks which flow through such rapidly growing cities as Delhi, Kolkata, Coimbatore, Hyderabad, Indore, Kanpur, Patna, Vadodara, and Varanasi. Approximately 40,500 ha of land is being irrigated with urban waste water by the side of the river Musi in Andhrapradesh, India (Hoek, 2006). It is has been carried out in different districts and municipalities of Tamil Nadu for assessing the suitability of sewage water to be used in agriculture. A survey reported that urban raw sewage suffered from poor water quality

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parameters and from heavy metal contents, being above permissible levels, particularly Cd, Cr, Pb and Ni (Chandrasekaran and Rajkannan, 2003).

A city with a population of 500,000 and water consumption of 200 L d<sup>-1</sup> per person would produce approximately 85,000 m<sup>3</sup> d<sup>-1</sup> of wastewater which could carefully irrigate an area of up to 6,000 ha, the fertilizer contribution of the effluent being N - 250 kg ha<sup>-1</sup> Year<sup>-1</sup>, P - 50 kg ha<sup>-1</sup> Year<sup>-1</sup>, K - 150 kg ha<sup>-1</sup> Year<sup>-1</sup> (FAO, 2004). Although lots of macro nutrients are added to soil, trace quantities of heavy metals also accumulate in the soil (Poornima, 2005). The more localized contamination found in urban environments is remediated through metal immobilization and mobilization that include bioremediation and chemical washing (Naidu, 2004). Removal of metals through chemical washing attracts research and commercial interest. However, very few works have been documented on heavy metal contamination of crops via sewage irrigation and none exist in relation to *Amaranthus cruentus* which is one of the most popular leafy vegetables in India. Because of the relative importance and toxicity of lead (Pb) and cadmium (Cd), even in trace concentrations, they were taken to be monitored in the present study.

Among many types of wastewater, foremost attention is being paid to sewage water due to its many attributes. Firstly, generation of sewage is inevitable in all towns and villages. Secondly, the impurity constituents of sewage are relatively less hazardous when compared with industrial effluents. Thirdly, the sewage carries considerable amounts of less complex organic matter along with essential nutrients required for plant growth. However, a prevention of bioavailability and transfer of toxic heavy metals into the food chain is the need of the hour. Therefore to address the above problems, the study was carried out in batch, pot and as well in the field.

## MATERIALS AND METHODS

Relevant materials and methods were employed to carry out batch, pot and field experiments. Field experiments were carried out on the sewage farm at Ukkadam, an intensively sewage drained area for over 65 years, located in Coimbatore peri-urban Tamil Nadu, India

### Urban Waste Water

Sewage effluent samples were collected from near the field experimental site. The samples were collected in polyethylene containers to be used in pot culture experiments at regular time intervals. The sub samples were stored following addition of toluene to arrest any microbial population and then stored at 4°C throughout the study. The samples were brought to room temperature prior to analysis. Initial physico-chemical parameters were analyzed by methods given by Standard Methods for the Examination of Water and Wastewater, (APHA 2005).

### Collection and Processing of Soil Samples

Soil samples were collected in polyethylene bags. The soil samples were shade dried and sieved through 2 mm sieves before being stored in other polyethylene bags for further analysis.

### Total Pb in Soil

The air dried and sieved soil samples were digested in aqua-regia (3:1, HCl: HNO<sub>3</sub>) until white precipitate appeared in the conical flask. Following a completion of digestion, the contents were transferred to 25 mL volumetric flasks, and then the volume made up to the mark. The supernatant was later passed through Whatman No 42 filter paper stored for later on metal analyses.

### Bio-available Fraction of Pb

The bioavailability of metals were studied as described by Shuman (1985). For water soluble and exchangeable metals, a batch experiment was conducted using polyethylene centrifuge tubes of 40 mL capacity. Soil solution was shaken in an end to end shaker at room temperature (25°C) for 6 hours for water soluble metals. A solution ratio of 1:2.5 (10 gram of soil to 25 mL of double distilled water) was employed. Following the tubes with the contents were centrifuged at 10,000 rpm for 10 minutes and then the supernatant filtered through Whatman No 42 filter paper. Following an extraction of the soil solution, 1M Ca(NO<sub>3</sub>)<sub>2</sub> was added to the remaining soil residue (at a ratio of 1:2.5), and kept for 4 hours in an end to end over shaker for exchangeable fraction of Pb. The rest of the extraction procedure followed was the same as that mentioned for water soluble fraction (Avudainayagam *et al.*, 2003)

### Batch Experiment

Batch experiments were conducted using centrifuge tubes of 40 mL capacity to find out the behavior of water soluble as well as exchangeable fractions of Pb in the soil. The chemical immobilizers namely FeSO<sub>4</sub> and CaSO<sub>4</sub> were added to the soil on a soil weight basis. The batch experiments were conducted in soil solutions of 1:2.5 (10 gm of soil to 25 mL of distilled water) ratios. The following chart depicts the treatments imposed for the batch experiment.

A kinetic experiment was also conducted to find out the releasing behavior of water soluble and exchangeable Pb from the sewage irrigated soil. The centrifuge tubes with their contents were shaken for various time periods of 0.5, 24, 48, and 168 hours. Following the respective shaking periods, the tubes were centrifuged with the contents processed for later analysis of water soluble and exchangeable Pb.

### Pot Culture Experiment

As for pot culture experiment, two doses of FeSO<sub>4</sub> viz. 1% (T<sub>2</sub>), and 2.5% (T<sub>3</sub>) were taken vs. control. Four crops were tested in the pot culture experiment. Two crops, namely *Amaranthus dubius* (C<sub>1</sub>) and *Beta vulgaris* var *bengalensis* (C<sub>2</sub>) are among green leafy vegetables and two flower crops namely *Celosia argentea* (C<sub>3</sub>) and *Tagetes erecta* (C<sub>4</sub>). Each treatment was replicated three times. A statistical design of Factorial and Completely Randomized Design (FCRD) was employed for the study.

### Collection and Processing of Plant Samples

Plant samples were collected from the pot culture and as well from field experiments at the end of the crop growing periods. The freshly collected crops were initially washed using tap water to remove any attaching soil particles. They were then washed, three washings with distilled water. The root, shoot and flower were separate and dried at 70°C for the coming steps. The dried plant

Treatment	Amount (gm kg <sup>-1</sup> soil )
Control (T <sub>1</sub> )	Soil alone
FeSO <sub>4</sub> 7H <sub>2</sub> O* 1% (T <sub>2</sub> )	0.10
FeSO <sub>4</sub> 7H <sub>2</sub> O* 2.5% (T <sub>3</sub> )	0.25
FeSO <sub>4</sub> 7H <sub>2</sub> O* 5% (T <sub>4</sub> )	0.50
CaSO <sub>4</sub> 2H <sub>2</sub> O (Lower requirement) (T <sub>5</sub> )	0.004
CaSO <sub>4</sub> 2H <sub>2</sub> O (Optimum requirement) (T <sub>6</sub> )	0.008
CaSO <sub>4</sub> 2H <sub>2</sub> O (Higher requirement) (T <sub>7</sub> )	0.013



samples were ground into powder using Willey mill for further analysis. A known weight of roots, shoots and flowers of the plant samples were taken to be digested with tri-acid mixture (9: 2: 1 of  $\text{HNO}_3$ :  $\text{H}_2\text{SO}_4$ :  $\text{HClO}_4$ ). The digested materials were filtered through Whatman No 42 filter paper and then made up to 25ml volume. The filtrates were used to measure for Pb through an Atomic Absorption Spectrometer (model Varian Spectra AA 100/200 FAAS).

### Field Experiment

Only the best performing flowering crop of *Tagetes erecta* alone was taken for study in the field conditions. In addition,  $\text{FeSO}_4$  doses were further decreased to 0.5 % ( $T_2$ ) and 1% ( $T_3$ ) along with control treatment ( $T_1$ ) for the field experiment. All the treatments were replicated seven times in a Randomized Block Design (RBD) of the statistical analysis.

### Statistical Analysis

The data recorded during the course of investigation was processed, tabulated and analyzed for statistical significance (Panse and Sukhatme, 1985).

## RESULTS AND DISCUSSION

### Initial Characterization of Soil and Sewage

The pH around neutral 7.57, along with EC ( $0.57 \text{ dS m}^{-1}$ ) of the soil fall in the limit prescribed for irrigation for raising agricultural crops. The soil is rich in organic carbon (3.27%). In contrast, sewage contains BOD ( $250$  to  $360 \text{ mg L}^{-1}$ ) and COD ( $490$ – $690 \text{ mg L}^{-1}$ ) which are beyond the maximum level for surface irrigation (Tables 1 and 2).

### Batch Experiment

In the batch experiments, analytical grade,  $\text{FeSO}_4$  and  $\text{CaSO}_4$  (Merck) were

**Table 1.** Physio-chemical characteristics of sewage irrigated soil.

Bulk density $\text{Mg m}^{-3}$	1.09
Particle density $\text{Mg m}^{-3}$	2.12
pH	7.57
EC $\text{dS m}^{-1}$	0.57
Organic carbon %	3.27
Total nitrogen %	0.12
$\text{KMnO}_4\text{-N kg ha}^{-1}$	255
Total phosphorus %	0.029
Olsen-P $\text{kg ha}^{-1}$	55
Total potassium %	0.15
Available potassium $\text{kg ha}^{-1}$	257
Total Ca %	1.79
Total Mg %	1.01
Total sodium %	0.12
Cation exchange capacity $\text{cmol (P}^+) \text{ kg}^{-1}$	12.6

**Table 2.** Chemical characteristics of sewage effluent.

pH	7.59
EC $\text{dS m}^{-1}$	1.21
Total dissolved solids $\text{mg L}^{-1}$	2450
Total suspended solids $\text{mg L}^{-1}$	2900
BOD $\text{mg L}^{-1}$	360
COD $\text{mg L}^{-1}$	690

used to immobilize the Pb in soil solution. Addition of  $\text{FeSO}_4$  decreased the soil pH significantly and the decrease in the trend was more with increase in the dose of  $\text{FeSO}_4$  (Table 3). Among the treatments, application of  $\text{FeSO}_4$  5% showed greater decrease in soil pH (from 7.54 to 5.55) within half an hour shaking period ( $T_4$ ). The decrease in pH was gradual until 24 hours shaking period, and due to soil buffering action pH might have been stabilized at 168 hours after the shaking was stopped. Similarly, an obvious increase in soil EC was also observed in  $\text{FeSO}_4$  and  $\text{CaSO}_4$  treated soils (Table 4). The increase in soil EC was more pronounced in  $\text{FeSO}_4$  than in  $\text{CaSO}_4$  treated soil. The increase was also more with increase in the dose of treatment. Among the doses, application of  $\text{FeSO}_4$  5% showed a greater increase in EC of the

**Table 3.** Effect of various amendments on soil pH.

Treatment	Shaking period				
	0.5 Hrs	24 Hrs	48 Hrs	168 Hrs	Mean
T <sub>1</sub>	7.54	7.54	7.54	7.54	7.54
T <sub>2</sub>	6.15	6.01	6.23	6.21	6.15
T <sub>3</sub>	5.70	5.60	6.03	6.13	5.87
T <sub>4</sub>	5.55	5.44	5.70	6.10	5.70
T <sub>5</sub>	7.00	6.71	6.00	6.51	6.56
T <sub>6</sub>	7.06	6.63	6.45	6.57	6.68
T <sub>7</sub>	7.12	6.62	6.61	6.57	6.73
Mean	6.59	6.37	6.36	6.52	
	SEd			CD (0.05)	
T	0.05			0.09	
D	0.03			0.07	
T× D	0.09			0.18	

soil (from 0.77 to 3.40 dS m<sup>-1</sup>) for a shaking period of 168 hours.

#### Effect of Various Amendments on Bioavailable Pb in Batch Experiment

In order to generate quicker results, batch experiments were conducted to select the suitable amendments for immobilization of Pb in soil solution. The amendment FeSO<sub>4</sub> was found to be more effective in the immobilization of bio-available fractions of Pb in the contaminated soil as compared with CaSO<sub>4</sub> treatment. Effective

immobilization of Pb was achieved when FeSO<sub>4</sub> added and mixed with the soil. Bioavailable fractions are water soluble and exchangeable ones, both of which got decreased to Below Detectable Levels (BDL) as a result of FeSO<sub>4</sub> application. This was observed for both doses of FeSO<sub>4</sub> namely 1% and 2.5% (Table 5). In CaSO<sub>4</sub> applied soil, the bio-available concentrations of Pb got below the detectable limits in the treatment of maximum CaSO<sub>4</sub> (T<sub>7</sub>) applied to soil. The two remaining doses of CaSO<sub>4</sub> (T<sub>5</sub> and T<sub>6</sub>) did not substantially influence the bioavailable Pb concentrations in soil.

**Table 4.** Effect of various amendments on soil EC (dS m<sup>-1</sup>).

Treatment	Shaking period				
	0.5 Hrs	24 Hrs	48 Hrs	168 Hrs	Mean
T <sub>1</sub> <sup>a</sup>	0.67	0.67	0.77	0.77	0.72
T <sub>2</sub> <sup>b</sup>	2.57	2.53	2.70	2.70	2.63
T <sub>3</sub> <sup>c</sup>	2.63	2.70	2.60	2.50	2.61
T <sub>4</sub> <sup>d</sup>	2.87	3.37	3.00	3.40	3.16
T <sub>5</sub> <sup>e</sup>	1.43	1.27	2.80	2.41	1.98
T <sub>6</sub> <sup>f</sup>	1.70	1.77	3.27	3.00	2.43
T <sub>7</sub> <sup>g</sup>	2.50	2.67	2.94	2.45	2.64
Mean	2.05	2.14	2.58	2.46	
	SEd			CD(0.05)	
T	0.03			0.07	
D	0.02			0.05	
T× D	0.07			0.13	

<sup>a</sup> Control; <sup>b</sup> 1% Ferrous sulphate; <sup>c</sup> 2.5% Ferrous sulphate; <sup>d</sup> 5.0% Ferrous sulphate; <sup>e</sup> Lower requirement of Calcium sulphate (0.05%); <sup>f</sup> Normal requirement of Calcium sulphate (0.10%); <sup>g</sup> Higher requirement of Calcium sulphate (0.15%).

**Table 5:** Effect various amendments on bio-available Pb (mg kg<sup>-1</sup>) in batch experiment.

Treatment	Water soluble	Exchangeable
Control (T <sub>1</sub> )	0.003	0.003
FeSO <sub>4</sub> 1% (T <sub>2</sub> )	BDL	BDL
FeSO <sub>4</sub> 2.5% (T <sub>3</sub> )	BDL	BDL
FeSO <sub>4</sub> 5% (T <sub>4</sub> )	BDL	BDL
CaSO <sub>4</sub> 0.0043 g 100 g <sup>-1</sup> soil (T <sub>5</sub> )	0.003	0.003
CaSO <sub>4</sub> 0.0087 g 100 g <sup>-1</sup> soil (T <sub>6</sub> )	0.003	0.003
CaSO <sub>4</sub> 0.0131 g 100 g <sup>-1</sup> soil (T <sub>7</sub> )	BDL	BDL
SEd <sup>a</sup>	0.0001	0.0001
CD (0.05) <sup>b</sup>	0.0002	0.0002

<sup>a</sup> Standard error of deviation, <sup>b</sup> Critical difference (95%).

Addition of FeSO<sub>4</sub> might have created formation of high iron oxide precipitates (Brown colour) in the soil solution and it was observed more in FeSO<sub>4</sub> treatment of 5%. Therefore the 5% application rate of FeSO<sub>4</sub> was not included in the pot culture experiment.

### Pot Culture Experiment

The pot culture experiment was carried out to study the uptake of trace metals by green leafy vegetables and by flower crop plants and the immobilization of Pb at lower doses of FeSO<sub>4</sub> (Table 6). Both fractions of water soluble as well as exchangeable were found to be lowered to below detectable levels by an application of various doses of FeSO<sub>4</sub>.

**Table 6.** Effects of FeSO<sub>4</sub> on aquaregia extractable Pb (mg kg<sup>-1</sup>) content in soil.

Crop	T <sub>1</sub>	T <sub>2</sub>	T <sub>3</sub>
C1	55.65	55.27	55.24
C2	55.55	55.36	55.07
C3	55.55	55.37	55.08
C4	55.65	55.43	55.14
Mean	55.60	55.36	55.13
		SEd	CD (0.05)
	C <sup>a</sup>	0.053	0.110
	T <sup>b</sup>	0.046	0.095
	C × T	0.092	0.191

<sup>a</sup> Crop, <sup>b</sup> Treatment.

### Effect of FeSO<sub>4</sub> on Aquaregia Extractable Pb (mg kg<sup>-1</sup>) in Post Harvested Pot Culture Soil

The aquaregia extractable Pb was slightly decreased in FeSO<sub>4</sub> applied soils in comparison with control. This decreasing trend was observed for the four crops irrespective of the crop type. The concentrations of Pb slightly decreased in FeSO<sub>4</sub> treated soil with it being higher in 2.5% amended (treated) pot (Table 7). The decrease might be due to adsorption of Pb into Fe oxides and also due to the uptake of the metal by crops in FeSO<sub>4</sub> treated soil.

Influence of FeSO<sub>4</sub> on Pb accumulation in Roots, Shoots and Flowers of Various Crop Plants (mg kg<sup>-1</sup>) in Pot Culture Experiment

Increasing the dose of FeSO<sub>4</sub> further, decreased the translocation of Pb from soil to roots, shoots and flowers. Irrespective of type of crop plant the accumulation of Pb decreased with FeSO<sub>4</sub> application. The pot culture experiment was carried out to see the effect of FeSO<sub>4</sub> on the immobilization of bioavailable Pb and its uptake by green leafy vegetable and flowering crops (Table 8). Among the crops, Pb uptake was low in flower crops in comparison with the green leafy vegetable crops, in the face of an application of FeSO<sub>4</sub>. Uptake was more in FeSO<sub>4</sub> treatment of 2.5% than in the 1%. Decrease in accumulation of Pb was observed (by order) more in roots of *T. erecta* > *A. dubius* > *B. bengalensis* > *C. argentea*.

**Table 7.** Effect of FeSO<sub>4</sub> on bioavailable Pb (mg kg<sup>-1</sup>) in pot culture experiment.

Crops	Water soluble fraction			Exchangeable fraction		
	T <sub>1</sub>	T <sub>2</sub>	T <sub>3</sub>	T <sub>1</sub>	T <sub>2</sub>	T <sub>3</sub>
C <sub>1</sub>	0.002	BDL	BDL	0.003	BDL	BDL
C <sub>2</sub>	0.003	BDL	BDL	0.002	BDL	BDL
C <sub>3</sub>	0.003	BDL	BDL	0.002	BDL	BDL
C <sub>4</sub>	0.003	BDL	BDL	0.003	BDL	BDL
		SEd	CD (0.05)		SEd	CD (0.05)
	C	0.0001	0.0003	C	0.0004	0.0007
	T	0.0001	0.0002	T	0.0003	0.0006
	C × T	0.0002	0.0005	C × T	0.0006	0.0013

**Table 8.** Accumulation of Pb in various crops (mg kg<sup>-1</sup>) in pot culture experiment.

Crop	Root			Shoot			Flower		
	T <sub>1</sub>	T <sub>2</sub>	T <sub>3</sub>	T <sub>1</sub>	T <sub>2</sub>	T <sub>3</sub>	T <sub>1</sub>	T <sub>2</sub>	T <sub>3</sub>
C <sub>1</sub>	3.54	0.44	0.25	4.57	0.75	0.45	-	-	-
C <sub>2</sub>	3.32	0.52	0.25	5.24	0.62	0.42	-	-	-
C <sub>3</sub>	2.68	0.35	0.32	3.34	0.52	0.25	1.58	0.47	0.12
C <sub>4</sub>	1.68	0.19	0.09	1.88	0.25	0.22	0.36	0.06	0.04
Mean	2.80	0.38	0.29	3.76	0.53	0.42	0.97	0.26	0.08
		SEd	CD		SEd	CD		SEd	CD
			(0.05)			(0.05)			(0.05)
	C	0.017	0.035	C	0.006	0.013	C	0.005	0.011
	T	0.014	0.030	T	0.005	0.011	T	0.006	0.013
	C×T	0.029	0.060	C×T	0.011	0.023	C×T	0.008	0.019

with the decrease being 94%, 92%, 92% and 88 % for the addition of FeSO<sub>4</sub> at 2.5 % level. The accumulation of Pb in the plants' shoots was in the order of *C. argenticus* > *B. bengalensis* > *A. dubius* > *T. erecta* with the accumulation percentages being 92%, 91%, 90% and 88% respectively for the addition of FeSO<sub>4</sub> at 2.5% level. Between the two flowering crops, accumulation of Pb was more decreased in *C. argenticus* than in *T. erecta* with the decrease in accumulation being 92% and 88%, respectively, as by addition of FeSO<sub>4</sub> at a 2.5% level.

### Field Experiments

The field experiments were conducted at Ukkadam sewage farm with *T. erecta* as which had performed well in the pot experiment, as the test crop. The dose of FeSO<sub>4</sub> was further decreased to 1.0 and 0.5 % for the field experiments to account for

the quantity and cost effectiveness. Similar to batch and pot cultures, field experiments also showed significant decreasing trend in Pb concentration and other chemical parameters in the soil solution. The value of pH decreased from 7.55 to 6.23 in surface soil and from 7.78 to 6.47 in sub surface soil (Table 9). EC increased from 0.57 to 5.61 and from 0.59 to 5.83 dS m<sup>-1</sup> in surface and subsurface soils respectively for FeSO<sub>4</sub> 1%.

### Effect of FeSO<sub>4</sub> on Bioavailable Fraction of Pb in Field Experiment Soils (mg kg<sup>-1</sup>)

Similar to batch and pot culture experiments, FeSO<sub>4</sub> application in the field experiment showed a significant decrease in bio-available fractions of Pb. Both fractions got diminished to below detectable levels in surface and subsurface

**Table 9.** Influence of amendment addition to soil on soil pH, EC (dS m<sup>-1</sup>) in the field experiment.

Treatment	Surface soil (0-15 cm)		Sub surface soil (15-30 cm)	
	pH	EC (dS m <sup>-1</sup> )	pH	EC (dS m <sup>-1</sup> )
T <sub>1</sub>	7.55	0.57	7.78	0.59
T <sub>2</sub>	6.74	5.63	6.95	5.84
T <sub>3</sub>	6.23	5.61	6.47	5.83
SEd	0.012	0.017	0.012	0.016
CD(0.05)	0.026	0.036	0.026	0.035

**Table 10.** Effects of FeSO<sub>4</sub> on bioavailable Pb (mg kg<sup>-1</sup>) in surface and sub surface soil.

Treatment	Surface soil (0-15 cm)		Sub surface soil (15-30 cm)	
	Water soluble	Exchangeable	Water soluble	Exchangeable
T <sub>1</sub>	0.002	0.005	0.002	0.005
T <sub>2</sub>	BDL	BDL	BDL	BDL
T <sub>3</sub>	BDL	BDL	BDL	BDL
SEd	0.0002	0.0002	0.0002	0.0002
CD(0.05)	0.0003	0.0003	0.0003	0.0003

**Table 11.** Effects of FeSO<sub>4</sub> amendment on aquaregia extractable Pb in the field experiment soil (mg kg<sup>-1</sup>).

Treatment	Surface soil (0-15 cm)	Sub surface soil (15-30 cm)
T <sub>1</sub>	110.32	109.44
T <sub>2</sub>	106.50	106.05
T <sub>3</sub>	104..30	103.89
SEd	0.0031	0.0031
CD(0.05)	0.0066	0.0066

**Table 12.** Influence of FeSO<sub>4</sub> on Pb accumulation in root, shoot and flower of *T. erecta* (mg kg<sup>-1</sup>).

Treatment	Root	Shoot	Flower
T <sub>1</sub>	2.312	1.312	0.032
T <sub>2</sub>	0.510	0.315	BDL
T <sub>3</sub>	0.330	0.222	BDL
SEd	0.0003	0.000	0.0001
CD(0.05)	0.0006	0.001	0.0002

soils for both doses of FeSO<sub>4</sub> applied. (Table10)

#### Effects of FeSO<sub>4</sub> on Aquaregia Extractable Pb in field Experiment Soils (mg kg<sup>-1</sup>)

The concentrations of aquaregia extractable Pb decreased by 3, 5, 3, and 6% respectively for surface and subsurface soils for (T<sub>3</sub>) treatment, this is due to addition of FeSO<sub>4</sub> and the interaction effects.

Influence of FeSO<sub>4</sub> on Pb Accumulation in Root, Shoot and Flowers of *T. erecta* (mg kg<sup>-1</sup>)

Within the crop, Pb was more accumulated in roots than in shoots. The plant uptake rate and further translocation of Pb in *T. erecta* was decreased more as a result of FeSO<sub>4</sub> application of 1% than 0.5%. Pb was not detected in flowers of *T. erecta* for any of the doses of FeSO<sub>4</sub> applied.

#### CONCLUSIONS

The results indicate that application of FeSO<sub>4</sub> greatly decreases the bioavailability



and uptake of water soluble and exchangeable Pb from soil solution. The decrease in soil pH might have favoured the adsorption sites and also adsorption onto the oxides of the added  $\text{FeSO}_4$  (Reddy and Patrick 1977, Ma and Uren 1997 and Hirsch and Banin 1990).

Application of  $\text{FeSO}_4$  0.5% is economically feasible and imparts more environmental benefits than the application of  $\text{FeSO}_4 > 0.5\%$ . Among the different doses of  $\text{FeSO}_4$ ,  $\text{FeSO}_4$  0.5% decreased heavy metal bioavailability considerably, at the same time field experiments revealed that *T. erecta* had low accumulation of Pb as compare with *C. argentia* and other green leafy vegetables. Benefit cost ratio for the treatment of  $\text{FeSO}_4$  0.5% was 1.05. The Benefit Cost ratio can be very much improved if includes the value of the environmental benefits. The chemical immobilization of trace metals in sewage contaminated sites will prevent the transfer of metals to the food chain and also stop the downward movement of the heavy metals into subsurface soil, which otherwise could cause the eventual contamination of ground water.

## REFERENCES

1. American Public Health Association (APHA), 2005. Eaton, A. D., Clesceri, L. S. Rice, R. W. and Greenberg, A. E. (Eds), Standard Methods for the Examination of Water and Wastewater 21st ed, American water works association; water pollution control federation, Washington, D.C. PP. 1368
2. APHA. 2005. *Standard Methods for the Examination of Water and Wastewater*. Amer. Pub. Health Assoc. and Amer. Water Works Assoc. and Amer. Water Pollution Conf. Federation, Broadway, New York. PP. 178:53-91
3. Avudainayagam, S., Megharaj, M., Owens, G., Kookana, R. S., Chittle Borough, D. and Naidu, R. 2003. *Chemistry of Chromium in Soils with Particular Emphasis on Tannery Waste Contaminated Sites: A Review*. Reviews in Environmental Contamination and Toxicology, 178, Springer-Verlag, New York.
4. Arti Bhatia, H. P. and Joshi, H. C. 2001. Use of Sewage as a Source of Plant Nutrient. *Fert. News.*, **46**: 55-58 and 61-64.
5. Chandrasekaran, N and Rajkannan, B. 2003. Nature and Characteristics of Sewage Effluents of Tamil Nadu. *Madras Agric. J.*, **90**: 124-128.
6. Elliott, H. A., Dempsey, B. A and Maille, P. J. 1990. Content and Fractionation of Heavy Metals in Water Treatment Sludges. *J. Environ. Qual.*, **19**: 330-334.
7. FAO, 2004. Wastewater Treatment and Use in Agriculture –Irrigation and Drainage Paper 47,2004, M. B. Pescod, M-56 ISBN 92-5-103135-5, PP. 10-11,
8. Hernandez-Allica, J., Garbisu, C., Barrutia, O and Becerril, J. M. 2007. EDTA Induces Heavy Metal Accumulation and Phytotoxicity in Cardoon Plants. *Environ. Exper. Bot.*, **60**: 26-32.
9. International Development Research Centre (IDRC). 2006. *Waste Water Use in Agriculture*. Ottawa, Canada.
10. Moolenaar, S. W. and Beltrami, P. 1998. Heavy Metal Balances of an Italian Soil as Affected by Sewage Sludge and Bordeaux Mixture Applications. *J. Environ. Qual.*, **27**: 828-835.
11. Naidu, R. 2004. United Approach to Assessing Contaminant Bioavailability in Soils. In: "Third International Workshop on Chemical Bioavailability in the Terrestrial Environment". Adelaide, South Australia, PP: 7-9.
12. Panse, V. G. and Sukhatme, P. V. 1985. *Statistical Methods for Agricultural Workers*. ICAR, New Delhi.
13. Poykio, R., Nurmesniemi, H. and Keisiki, R. L. 2007. Environmental Risk Assessment of Heavy Metal Extractability in a Biosludge from of the Biological Waste Water Treatment Plant of a Pulp and Pear Mill. *J. Environ. Monitoring Assess.*, **128**: 153-164.
14. Poornima, K., 2005. Immobilization and Uptake Studies in Heavy Metals in Sewage Contaminated Soil. MSc. (Ag.) Thesis Submitted to TNAU, Coimbatore.
15. Raschid-Sally, L., Tuan, D. D., Abayawardana, S. 2004. National Assessments on Wastewater Use in Agriculture and an Emerging Typology: The Vietnam Case Study. In: "Use in Irrigated Agriculture: Confronting the Livelihood and Environmental Realities" (Eds.): Scott, C. A., Aruqui, N. I., Raschid-Sally, L. Wastewater Commonwealth Agricultural Bureau International, Orient-Longman, and International Development Research Centre, Ottawa, Canada. PP. 81-90



16. Reddy, C. N. and Patrick, W. H. 1977. Effect of Redox Potential and pH on the Uptake of Cadmium and Lead by Rice Plants. *J. Environ. Qual.*, **6**: 259-262
17. Scott, C. A., Faruqui, N. I., Raschid-Sally, L., 2004. Wastewater Use in Irrigated Agriculture: Management Challenges in Developing Countries. In: "Wastewater Use in Irrigated Agriculture" (Eds.): Scott, C. A., Faruqui, N. I., Raschid-Sally, L., CABI Publishing, UK.
18. Strauss, M., Blumenthal, J. U. 1990. Use of Human Waste Use in Agriculture and Aquaculture-utilization Practices and Health perspectives. IRCWD (International Reference Center for Waste Disposal) Duebendorf, Switzerland. Report No. 09/90.
19. Udom, B. E., Mbagwu, J. S. C., Adesodun, J. K. and Agbim, M. N. 2004. Distribution of Zinc, Copper, Cadmium and Lead in a Disposal of Sewage Sludge. *Environ. Int.*, **30**: 467-470.
20. Van der Hoek, W. 2004. A Framework for a Global Assessment of the Extent of Wastewater Irrigation: The Need for a Common Wastewater Typology. In: "Wastewater Use in Irrigated Agriculture: Confronting the Livelihood and Environmental Realities" (Eds.): Scott, C. A.; Faruqui, N. I.; Raschid-Sally, L., Wallingford, UK: CABI Publishing. PP. 11-24.
21. Wang, P. E., Qu Li, Z. and. Shuman, L. M. 1997. Fractions and Availability of Nickel in Loessial Soil Amended with a Sewage Sludge. *J. Environ. Qual.*, **26**: 795-801.

### تثبیت شیمیایی سرب (Pb) در خاک آبیاری شده با فاضلاب به مدت طولانی

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#### چکیده

پساب فاضلاب حاوی مقادیر قابل توجهی نیتروژن (۰/۰۹ تا ۰/۱۶٪)، فسفر (۰/۲۳ تا ۰/۲۹٪) و پتاسیم (۰/۱۱ تا ۰/۱۸٪) بوده و همچنین غنی از کربن آلی (۳/۲۷٪) می باشد. در هنگام آبیاری با این آب، فلزات سمی نیز در منطقه انباشته می شوند. برای جلوگیری سریع از انتقال فلز کمیاب سمی سرب از خاک آلوده به محلول خاک و ورود آن به چرخه غذا می توان از  $\text{CaSO}_4$  و  $\text{FeSO}_4 \cdot 7\text{H}_2\text{O}$  استفاده نمود. در این مطالعه، کارایی تثبیت  $\text{CaSO}_4$  و  $\text{FeSO}_4 \cdot 7\text{H}_2\text{O}$  در آزمایشگاه و مزرعه مورد بررسی قرار گرفت. در آزمایش ناپیوسته، استفاده از  $\text{FeSO}_4$  به میزان ۵٪ بر مبنای وزن خاک باعث کاهش شدید pH خاک از ۷/۵۴ به ۵/۵۵ در مدت نیم ساعت تکان دادن و در نتیجه کاهش صددرصدی در غلظت سرب قابل تبادل و محلول در آب گردید. استفاده از  $\text{CaSO}_4$  نیز موجب کاهش pH و همچنین غلظت سرب در محلول خاک گردید. آزمایش های کشت گلدانی با استفاده از دو نوع گیاه (گلدار و برگری) نیز کاهش دسترسی زیستی سرب را در خاک های تیمار شده با  $\text{FeSO}_4$  نشان داد.